

- What is the solubility of $\text{Cu}(\text{OH})_2$ in mol L^{-1} ? $K_{\text{sp}}(\text{Cu}(\text{OH})_2)$ is 1.6×10^{-19} at 25°C .

Marks
6

Answer:

The overall formation constant for $[\text{Cu}(\text{NH}_3)_4]^{2+}$ is 1.0×10^{13} . Write the equation for the reaction of Cu^{2+} ions with excess ammonia solution.

Calculate the value of the equilibrium constant for the following reaction.



Answer:

Would you expect $\text{Cu}(\text{OH})_2(\text{s})$ to dissolve in 1 M NH_3 solution? Briefly explain your answer.