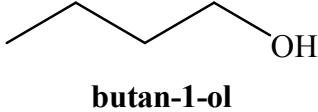
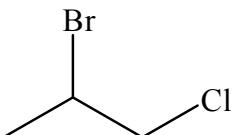
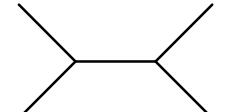
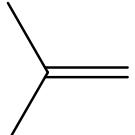
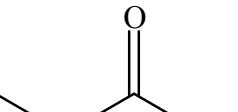
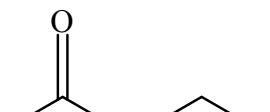
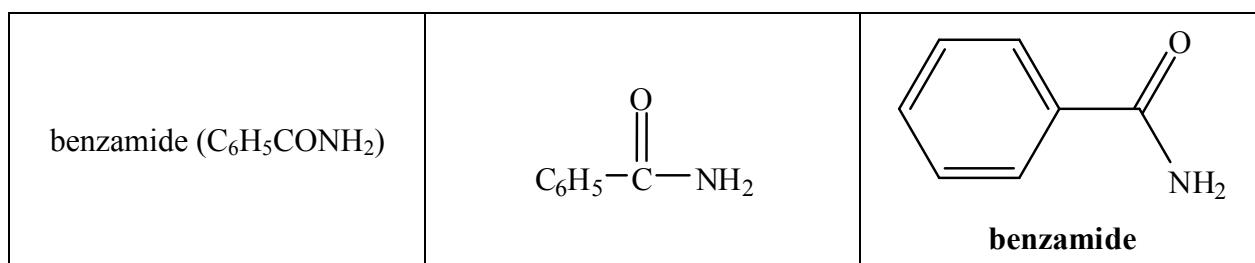


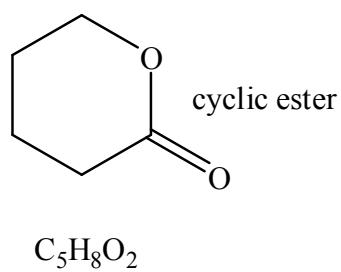
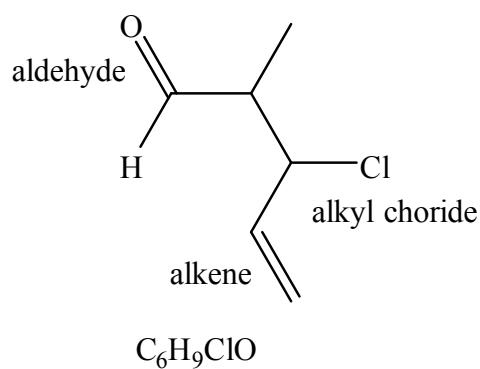
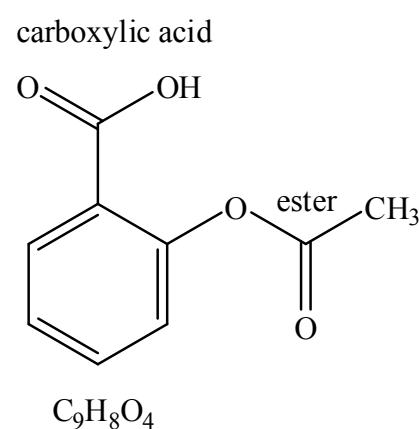
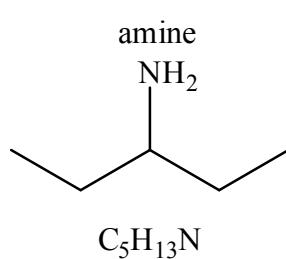
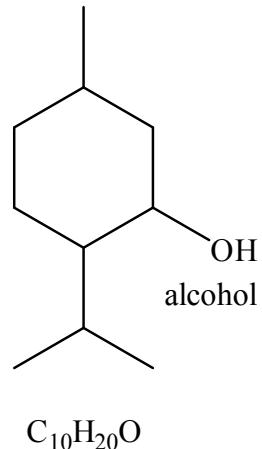
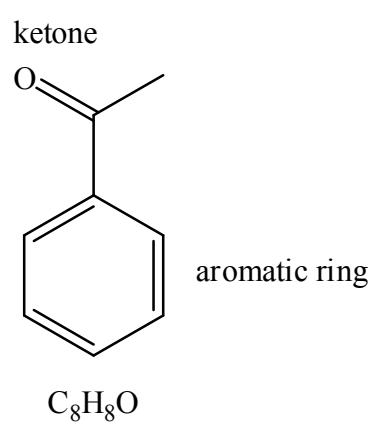
CHEM1102 Answers to Problem Sheet 1 (Week 1)

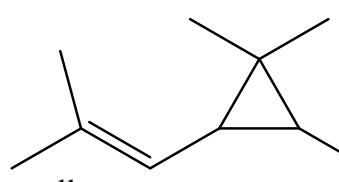
1.

Compound	Condensed structural formula	Stick representation
$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$	$\text{CH}_3-\text{CH}_2-\text{CH}_2-\text{CH}_2\cdot\text{OH}$	 butan-1-ol
$\text{CH}_3\text{CHBrCH}_2\text{Cl}$	$\begin{array}{c} \text{Br} \\ \\ \text{CH}_3-\text{CH}-\text{CH}_2-\text{Cl} \end{array}$	 2-bromo-1-chloropropane
$(\text{CH}_3)_2\text{CHCH}(\text{CH}_3)_2$	$\begin{array}{ccccc} & \text{H}_3\text{C} & & \text{CH}_3 & \\ & & & & \\ & \text{CH} & - & \text{CH} & \\ & & & & \\ & \text{H}_3\text{C} & & \text{CH}_3 & \end{array}$	 2,3-dimethylbutane
$(\text{CH}_3)_2\text{CCH}_2$	$\begin{array}{c} \text{H}_3\text{C} \\ \\ \text{H}_3\text{C}-\text{C}=\text{CH}_2 \\ \\ \text{H}_3\text{C} \end{array}$	 2-methylprop-1-ene
$\text{CH}_3\text{CH}_2\text{COOH}$	$\text{CH}_3-\text{CH}_2-\overset{\text{O}}{\underset{\parallel}{\text{C}}}-\text{OH}$	 propionic acid
$\text{CH}_3\text{COOCH}_2\text{CH}_3$	$\text{CH}_3-\overset{\text{O}}{\underset{\parallel}{\text{C}}}-\text{O}-\text{CH}_2-\text{CH}_3$	 ethyl acetate
$\text{CH}_3\text{COOCOCH}_3$	$\text{CH}_3-\overset{\text{O}}{\underset{\parallel}{\text{C}}}-\text{O}-\overset{\text{O}}{\underset{\parallel}{\text{C}}}-\text{CH}_3$	 acetic anhydride

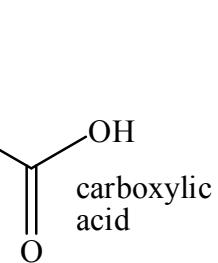
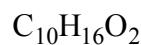


2.

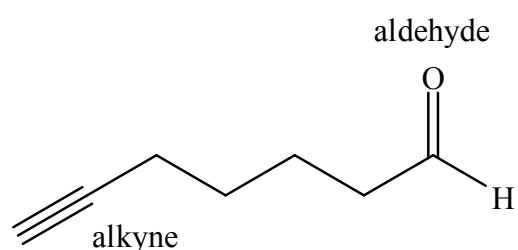




alkene



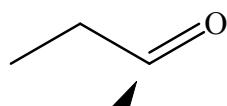
carboxylic
acid



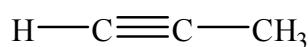
alkyne



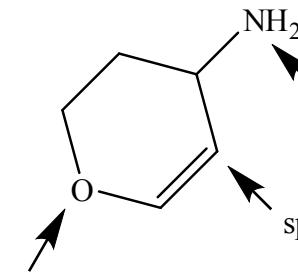
3.



sp^2 (trigonal planar carbon
bond angles $\sim 120^\circ$)

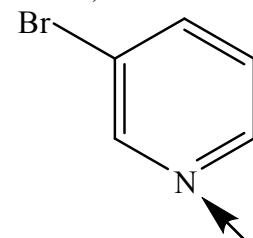


sp (linear
bond angles = 180°)

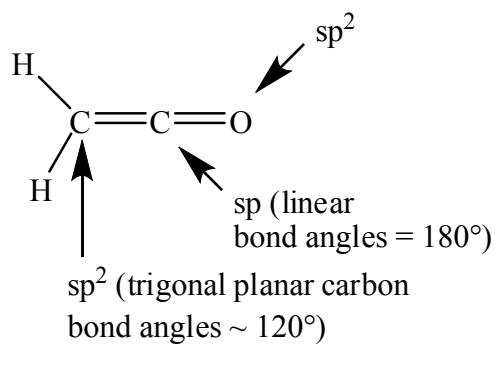


sp^3 (tetrahedral
bond angles $\sim 109^\circ$)

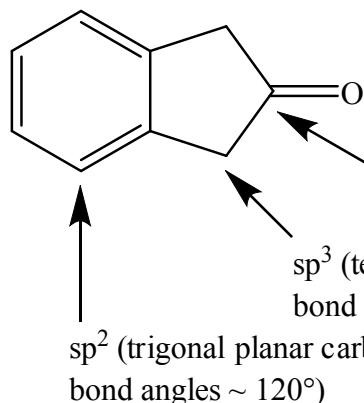
sp^2 (trigonal bond angles $\sim 120^\circ$)



sp^2 (trigonal planar
bond angles $\sim 120^\circ$)



sp^2 (trigonal planar carbon
bond angles $\sim 120^\circ$)



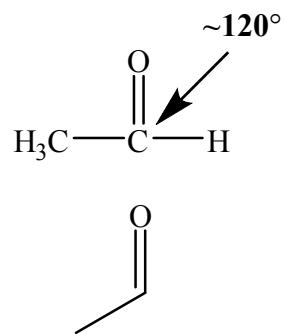
sp^2 (trigonal planar carbon
bond angles $\sim 120^\circ$)

sp^3 (tetrahedral carbon
bond angles $\sim 109^\circ$)

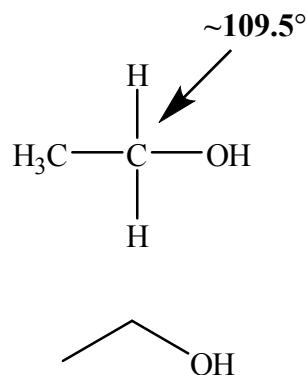
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4.

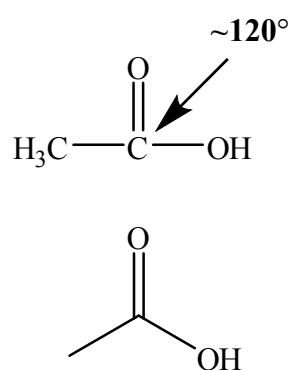
(a)



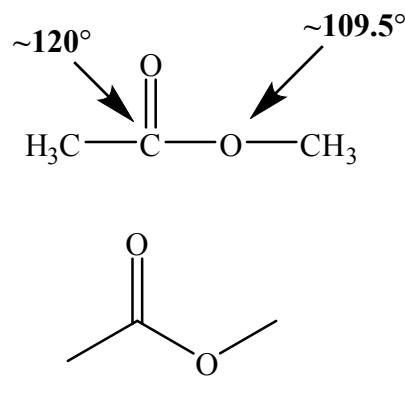
(b)



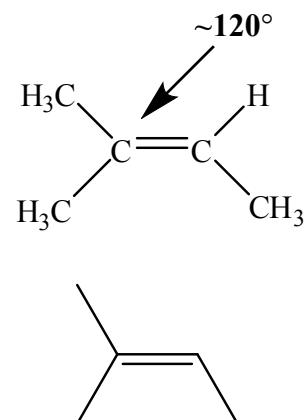
(c)



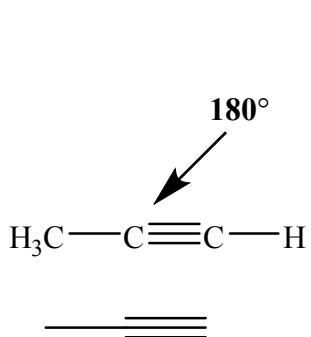
(d)



(e)



(f)



(g)

